

عنوان مقاله:

Application of the Specific Ion Interaction Theory to the Ionic Strength Dependence of Aqueous Dissociation Constants of Naringenin in NaClO₄ Solutions

محل انتشار:

سومین همایش ملی کاربردهای شیمی در فناوریهای نوین (سال: 1392)

تعداد صفحات اصل مقاله: 6

نویسندگان:

Morteza Jabbari - *Department of Chemistry, Neyshabour Branch, Islamic Azad University, Neyshabour, Iran*

Raheleh Zhiani - *Department of Chemistry, Neyshabour Branch, Islamic Azad University, Neyshabour, Iran*

Ali Farajtabar - *Department of Chemistry, Jouybar Branch, Islamic Azad University, Jouybar, Iran*

خلاصه مقاله:

The aqueous dissociation constants of naringenin were obtained from potentiometric and spectrophotometric measurements at different ionic strength values and 25 °C. Ionic strength of solution was supplied by NaClO₄ in the range of 0.10-4.0 mol dm⁻³. The dependence of protonation constant on ionic strength was described by Specific Ion Interaction Theory (SIT), and the specific interaction parameters of the ionic species were determined. The protonation constant values were also calculated at infinite dilution.

کلمات کلیدی:

Protonation constant, Naringenin, SIT model, Ionic strength

لینک ثابت مقاله در پایگاه سیویلیکا:

<https://civilica.com/doc/233179>

